





Impact of calcium and copper co-doping on the oxygen transport of layered nickelates: a case study of $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ and a comparative analysis

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Abstract

This study aims at investigating oxygen content and diffusion in the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ ($y = 0.0-0.4$) series. Single-phased materials were obtained via nitrate combustion using glycerol as a fuel. The materials exhibit an orthorhombic structure, which correlates well with the absolute oxygen content in the samples in the range of 4.15–3.97. The oxygen mobility in the samples was investigated by a temperature-programmed isotope exchange of oxygen (TPIE) with C^{18}O_2 in a flow reactor. The obtained curves, comprising few extrema, were fitted using a mathematical model that includes two or three distinct oxygen forms, each characterized by a specific diffusion coefficient value. The oxygen tracer diffusion coefficient values and the contribution of fast oxygen forms tend to decrease with increasing Cu content (except for $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.8}\text{Cu}_{0.2}\text{O}_{4+\delta}$), which can be explained by the variation of highly mobile interstitial oxygen content since the effective activation energy values remains the same within the calculation error. Comparative analysis shows that this behavior is similar to the trends observed and discussed for Ca- and Cu-doped La and Nd nickelates. Nevertheless, the highest D^* value for O_{fast} was demonstrated for the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.8}\text{Cu}_{0.2}\text{O}_{4+\delta}$ sample ($9.4 \cdot 10^{-8} \text{ cm}^2/\text{s}$ at $700 \text{ }^\circ\text{C}$), which correlates well with superior electrochemical properties of the electrodes on its basis.

Keywords

solid oxide fuel cells
solid oxide electrolyzers
layered nickelates
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Key findings

- Oxygen tracer diffusion coefficient values tend to decrease with increasing Cu content.
- There is a nonuniformity in oxygen mobility expressed in 2 or 3 form of oxygen.
- The trends in the oxygen transport properties are similar to those for the Ca- and Cu-doped La and Nd nickelates.

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1. Introduction

The $\text{Ln}_2\text{NiO}_{4+\delta}$ ($\text{Ln} = \text{La}, \text{Pr}, \text{Nd}$) phases, as representatives of the first-order Ruddlesden–Popper (RP) phases, have been extensively studied in the last two decades as air electrodes for solid oxide fuel cells (SOFCs) [1–8], solid oxide electrolysis cells (SOECs) [9–13], oxygen separation membranes [14–19], and catalysts for hydrocarbon fuel

conversion into syngas [20, 21] and other reactions [22, 23]. The variety of applications has been attributed to several factors, including a moderate thermal expansion coefficient, stability under operating conditions, high catalytic and electrocatalytic activity in the respective reactions, low polarization resistance, and high oxygen permeability. These properties have been provided by the layered structure, high mixed ionic (O^{2-} and occasionally H^+)-electronic

conductivity, and ability to exsolve metal nanoparticles under reducing conditions [20, 23].

The high oxygen ionic conductivity characteristics of RP phases are due to the cooperative mechanism of oxygen migration via interstitialcy movement involving both regular and interstitial oxygen [24–27]. The interstitial oxygen of the rock salt layer is known to be highly mobile and capable of accumulating to a high extent. This mechanism results in high oxygen self-diffusion coefficient values (up to $\sim 10^{-8}$ – 10^{-7} cm²/s at 700 °C [28,29]), which are attractive for electrochemical and catalytic applications [1,7,14]. At the same time, the total conductivity of RP phases is moderate and exhibits a bell-shaped dependence on temperature, reaching a maximum in the 400–550 °C range, followed by a decrease at higher temperatures due to the reduction in charge carriers caused by oxygen release from the layered structure [1, 30]. It is essential to enhance the maximal conductivity values, which are typically less than 10² S cm⁻¹ for undoped Ln₂NiO_{4+δ}, with the objective of facilitating more efficient electrochemical and catalytic applications. A number of studies demonstrated that doping Ln-site with alkaline earth metals (AEMs) enhances the total conductivity, particularly when Ca is employed as a dopant (in comparison with oxides with equal AEM dopant concentrations) [31–33]. Nevertheless, AEM doping resulted in a reduction in oxygen mobility, attributed to a decrease in the interstitial oxygen content and steric hindrances for the cooperative mechanism [29, 34–36]. Additionally, AEM doping was shown to increase the thermal expansion coefficient [1, 7].

Doping at the Ni-site is another approach to modify the functional properties of RP phases. Doping with cations such as Mg, Mn, Fe, Co, Cu, Ga or their combinations can result in modifying the thermal and chemical expansion behavior of the oxides and compensating for the negative effects or enhancing the positive influence of Ln-site doping [37–43]. Ni-site doping can either increase or decrease the oxygen mobility and may sometimes have an even more pronounced influence than Ln-site doping [39, 44–48].

Typically, the doping of Ln₂NiO_{4+δ} at the Ni-site with Cu results in a decrease in total conductivity; however, at certain concentrations, it can improve the thermomechanical compatibility of RP phases with solid-state electrolytes and increase the electrochemical activity of the electrodes [40, 45, 49–51]. Nevertheless, the information on the effects of doping with Cu on the oxygen transport properties is contradictory. Doping Ln nickelates with Cu resulted in either an increase or a decrease in the oxygen mobility due to the effects of the energy barrier for interstitial oxygen migration, interstitial oxygen content, grain growth features, changes in the (Ni,Cu)–O bond length, and other factors [46, 52–55].

This work continues the series of systematic studies concerning the structure, functional properties and electrochemical applications of Ln₂NiO_{4+δ} (Ln = Nd, La, Pr), double doped with Ca and Cu at the Ln and Ni sites,

respectively [40, 44, 47, 56–58]. The novelty of this work is a comprehensive study of oxygen transport characteristics of the Pr_{1.6}Ca_{0.4}Ni_{1-y}Cu_yO_{4+δ} series, which is performed for the first time. This paper also provides a summary of the structural properties, oxygen content and other characteristics of the materials of the Ln_{1-x}Ca_xNi_{1-y}Cu_yO_{4+δ} series, as well as a comparative analysis of their functional properties and electrochemical characteristics. It was demonstrated that Pr-based materials exhibit superior electrochemical performance in this series. For example, the polarization resistance (R_p) of the Pr_{1.6}Ca_{0.4}Ni_{0.8}Cu_{0.2}O_{4+δ} electrode was found to be in the range of 0.19–0.27 Ω cm² (at 700 °C) depending on the electrolyte substrate (Ce_{0.8}Sm_{0.2}O_{2-δ}, doped BaCe(Zr)O_{3-δ}) [57] compared with those obtained for Nd-based (2.73–4.79 Ω cm² at 700 °C [47] and La-based (1.95 Ω cm² at 700 °C [40]) electrodes. Given that RP phases are mixed ion- and electron-conducting (MIEC) materials and that their electrochemical properties are strongly dependent on their oxygen transport properties [59], this study aims to investigate the oxygen content and diffusion in the Pr_{1.6}Ca_{0.4}Ni_{1-y}Cu_yO_{4+δ} ($y = 0.0$ – 0.4) series. The data obtained were employed to determine the correlations between the oxygen diffusion kinetics and electrochemical properties, as well as to compare them with undoped, Ca-doped and Ca/Cu-doped Ln₂NiO_{4+δ}.

2. Materials and methods

Materials of the Pr_{1.6}Ca_{0.4}Ni_{1-y}Cu_yO_{4+δ} ($y = 0.0$ – 0.4 , PCNO, PCNCO1–PCNCO4) series were synthesized by nitrate combustion using glycerol as a fuel, with high purity reagents, namely, Pr₆O₁₁ (98% purity), CaCO₃ ($\geq 97\%$ purity), Ni(CH₃COO)₂·4H₂O (98% purity), and CuO ($\geq 99\%$ purity). The oxides were pre-annealed for 4 h at 600 °C to remove adsorbed moisture and CO₂, then taken in stoichiometric amounts and dissolved in nitric acid ($\geq 65\%$ purity). The salt Ni(CH₃COO)₂·4H₂O was dissolved in distilled water. The resulting solutions were then combined, and glycerol (99.6% purity) was subsequently added to the mixture to serve the dual functions of a chelating agent and an organic fuel. The required amount of glycerol was determined from the redox reaction equation, wherein the nitrate ion serves as the oxidizing agent.

The resulting precursor was heated on a hot plate until the water had evaporated completely, thereby initiating the combustion process. The resulting powder was then subjected to annealing at 900 °C for 12 h to remove any residual organic matter. The synthesis of the materials was conducted in two stages, at 1000 °C for 12 h and at 1100 °C for 12 h, with intermediate grinding occurring between the stages. Details of the synthesis procedure are given in [57].

The phase purity of the Pr_{1.6}Ca_{0.4}Ni_{1-y}Cu_yO_{4+δ} ($y = 0.0$ – 0.4) materials was determined after the final step by powder X-ray diffraction (XRD) analysis using a Shimadzu XRD-7000 diffractometer (Japan) with a graphite monochromator with Cu K α 1 radiation at room temperature in air. The

XRD patterns were obtained over the angle range of $10 < 2\theta < 90$ with a step size of 0.02° and an exposure time of 5 s at each point.

Given that the combustion products were typically highly agglomerated, the obtained powders were thoroughly dispersed in an agate mortar with the addition of ethanol to achieve a specific surface area (S_{BET}) of no less than $1 \text{ m}^2 \text{ g}^{-1}$. The Brunauer-Emmett-Teller (BET) method was employed to determine S_{BET} values, utilizing a SORBI N.4.1 analyzer (Meta, Russia).

The temperature behavior of the oxygen overstoichiometry, δ , in the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ ($y = 0.0-0.4$) materials and the absolute oxygen content were determined via thermogravimetric analysis (TGA) by heating the samples in air and in a H_2+Ar gas mixture (volume ratio of 1:1), respectively. The weight loss measurements were conducted using an STA 449 F3 Jupiter thermal analyzer (Netzsch, Germany). Prior to measurements, the oxygen content in the powder samples was equilibrated in dry air at 1100°C for 1 h inside the device with slow heating and cooling. The samples were then heated in air to 1000°C at a rate of $2^\circ\text{C}/\text{min}$ to measure the weight loss, from which the temperature dependence of the change in δ was subsequently determined. The powders were cooled and then heated to 1000°C in a H_2+Ar gas mixture (volume ratio of 1:1) until complete reduction. On the basis of the results of both experiments, the δ values at room temperature and 700°C were determined. The experimental procedure was established based on the data obtained for the Cu-doped PNO (Supplementary data, Figure S1).

The oxygen mobility of the samples was investigated via the temperature-programmed isotope exchange of oxygen (TPIE) with C^{18}O_2 in a flow reactor. The samples sintered at 1100°C (Table 1, 50 mg weight) were pelletized under a pressure of 400 MPa, then ground and sieved in order to obtain 0.25–0.5 mm fraction. The true density of the samples was measured via hydrostatic weighting. The bulk density, which is the geometric density of the fraction, was calculated by dividing the mass of the fraction by the volume which it occupied.

The samples were loaded into a quartz tubular reactor (inner diameter of 3 mm). The samples were pretreated in a flow of $\text{He} + 1\% \text{ O}_2$ (flow rate of 25 ml/min) at 700°C for 0.5 h and then cooled. The TPIE experiments were performed in $\text{He} + 1\% \text{ C}^{18}\text{O}_2 + 1\% \text{ Ar}$ (flow rate of 25 ml/min)

while heating from 50°C to 800°C with a ramp rate of $5^\circ\text{C}/\text{min}$. The gas phase composition at the reactor outlet was analyzed by a QMS 200 (Stanford Research Systems, USA) mass spectrometer. The ^{18}O atomic fraction (α) and $\text{C}^{16}\text{O}^{18}\text{O}$ molecular fraction (f_{16-18}) in gas phase dependences were used for calculations of isotope exchange kinetic parameters via mathematical modeling.

3. Results and Discussion

3.1. Structure

The analysis of the X-ray diffraction data of the as-synthesized $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ powder samples reveals the formation of homogeneous complex oxide phases across the entire range of dopant concentrations from $y = 0.0-0.4$ (PCNO, PCNCO01–PCNCO04). The XRD patterns are presented in Figure 1. All of the obtained materials possess an orthorhombic structure with the $Bmab$ space group. The refinement of the unit cell parameters was carried out using the Rietveld full-profile analysis method using the FullProf Suite software (presented in Table 2). The theoretical density of the powders, ρ_{XRD} , was calculated as follows (Equation 1) [60]:

$$\rho_{\text{XRD}} = \frac{MZ}{N_A V} \quad (1)$$

where M is the molar mass, Z is the number of formula units in each unit cell (equal to 4 in this case), V is the unit cell volume, and N_A is the Avogadro number.

The unit cell parameter c increases with increasing Cu content in the series, whereas the a and b parameters decrease, which can be related to the tendency of Jahn-Teller Cu^{2+} ions to form elongated CuO_6 octahedra [61, 62]. These multidirectional changes result in an increase in the unit cell volume.

Table 1 Parameters of the samples used for the TPIE C^{18}O_2 studies.

Sample	S_{BET} , [m^2/g]	True density, [g/cm^3]	Bulk density, [g/cm^3]
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{NiO}_{4+\delta}$	1.5	5.76	2.07
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.9}\text{Cu}_{0.1}\text{O}_{4+\delta}$	2.1	5.96	2.31
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.8}\text{Cu}_{0.2}\text{O}_{4+\delta}$	1.3	5.58	1.98
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.7}\text{Cu}_{0.3}\text{O}_{4+\delta}$	1.0	5.62	2.11
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.6}\text{Cu}_{0.4}\text{O}_{4+\delta}$	1.2	5.83	2.09

Table 2 Goldschmidt tolerance factor (t), structure parameters and oxygen content for the PCNCO series.

y	0.0	0.1	0.2	0.3	0.4
Sample designation	PCNO	PCNCO01	PCNCO02	PCNCO03	PCNCO04
t	0.899	0.886	0.881	0.880	0.877
a , [Å]	5.3537(1)	5.3441(2)	5.3340(3)	5.3270(1)	5.3220(2)
b , [Å]	5.3726(2)	5.3686(3)	5.3626(3)	5.3580(1)	5.3560(2)
c , [Å]	12.3706(3)	12.4166(3)	12.4659(3)	12.5051(4)	12.5521(4)
V , [Å ³]	355.82(2)	356.24(1)	356.58(1)	356.98(2)	357.8(2)
ρ_{XRD} , [g/cm ³]	6.803	6.798	6.801	6.802	6.796
$4+\delta$	4.12	4.09	4.02	3.97	4.00

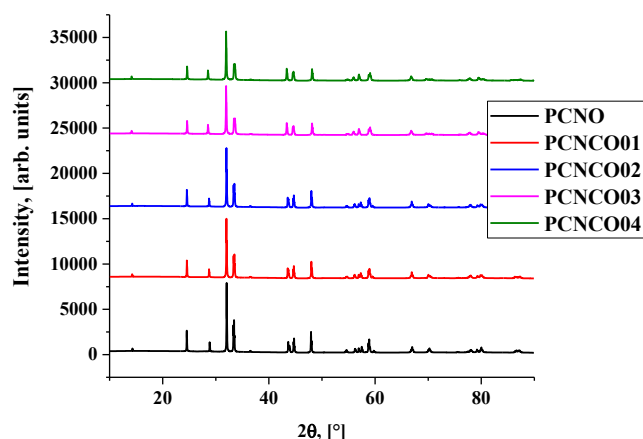


Figure 1 XRD patterns obtained at room temperature for the as-synthesized PCNCO materials.

In our recent studies, extended data on the structure for $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ ($y = 0.0-0.4$) [57], $\text{Nd}_{1.6}\text{Ca}_{0.4}\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ ($y = 0.0-0.4$) [47,56] and $\text{La}_{1.7}\text{Ca}_{0.3}\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ ($y = 0.0-0.4$) [40] materials, including calculated bond lengths, were presented. Table S1 provides a summary of the structural data for the initial, middle and final compositions of the series of materials ($y = 0, 0.2$ and 0.4 (0.3 in the case of the NCNCO series), PCNO, PCNCO02, PCNCO04, NCNO, NCNCO02, NCNCO03, LCNO, LCNCO02, and LCNCO04, respectively). It can be seen that the tolerance factor, t , calculated according to Equation S1, decreases slightly with increasing Cu concentration and is in the range of $0.87 < t < 0.99$, which is typical of the $\text{K}_2\text{NiF}_4(T)$ -type structure [63]. A decrease in t results from the replacement of Ni^{2+} by larger cations of Cu^{2+} in the series ($r_{\text{Ni}^{2+_{\text{VI}}}} = 0.69 \text{ \AA}$, $r_{\text{Ni}^{3+_{\text{VI}}(\text{LS})}} = 0.56 \text{ \AA}$, $r_{\text{Ni}^{3+_{\text{VI}}(\text{HS})}} = 0.59 \text{ \AA}$, $r_{\text{Cu}^{2+}} = 0.73 \text{ \AA}$ according to Shannon [64]), which causes an expansion of the $[\text{Ni}/\text{CuO}_2]^{2-}$ layers, thus destabilizing the structure. Nevertheless, single-phase materials are obtained over the whole concentration range up to $y = 0.4$ for the Pr/Ca and La/Ca series, confirming the higher solubility of Cu cations and their structures than those of Nd/Ca ($y = 0.35$) [47].

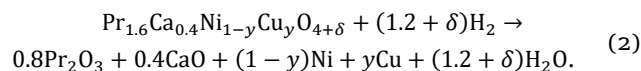
According to the data of the high-temperature XRD study [47,57], the PCNCO and NCNCO materials exhibit an orthorhombic structure at room temperature with the phase transition to a tetragonal structure at temperatures of approximately 180–200 and 300–400 °C, respectively (Table S1).

In Ca-doped layered nickelates, cations occupying the Ln/Ca sites coordinate with four O1 and five O2 ions, whereas in the first coordination sphere of the cations in the Ni sites there are four O1 ions and two O2 ions [65]. Copper cations are able to form elongated CuO_6 octahedra in place of more regular NiO_6 octahedra. This can be explained by the preference of Cu^{2+} cations for square planar

coordination [66], resulting in an elongated Ni/Cu–O2 bond length and a compressed Ni/Cu–O1 bond length, which is observed in the whole LnCNCO series (Table 2). The replacement of Ni^{2+} by Cu^{2+} results in an increase in the Ln/Ca–Ln/Ca, Ln/Ca–O1 and Ln/Ca–O2x1 bond lengths, whereas the Ln/Ca–O2x4 bond length decreases. The c unit cell parameter value, which is the sum of doubled Ln/Ca–O2x1, doubled Ni/Cu–O2, and Ln/Ca–Ln/Ca distances, increases with increasing Cu doping. The Ni/Cu–O1 and Ln/Ca–O2x4 bond lengths, corresponding to the a and b parameters, both decrease with increasing doping. Since the Ln/Ca–O2 distance corresponds to the thickness of the $\text{Ln}(\text{Ca})_2\text{O}_2$ layer, Cu doping may result in faster oxygen migration in the LnCNCO structure.

3.2. Oxygen content

The oxygen content in the samples was calculated according to the reaction of the material reduction in a hydrogen-containing atmosphere. For example, for the PCNCO series it can be written as follows (Equation 2):



The δ value in the material can be calculated from the change in the mass of the sample in the reducing atmosphere as follows:

$$\delta = \frac{-M_{\text{ox}}}{M_{\text{O}}} + \frac{m_0 M_{\text{red}}}{M_{\text{O}} m_{\text{red}}}, \quad (3)$$

where m_0 and m_{red} correspond to the mass of the sample before and after reduction, respectively, g ; M_{ox} is the molar mass of the initial oxide, $g \text{ mol}^{-1}$; M_{red} is the total molar mass of the reduction products, considering the coefficients in Equation (2), $g \text{ mol}^{-1}$; and M_{O} is the molar mass of the oxygen atom, $g \text{ mol}^{-1}$.

At the conventional operation temperature of IT-SOFCs (700 °C), the oxygen content is calculated from the TGA data obtained during heating in air. The resulting dependences for the PCNCO, NCNCO and LCNCO series are presented in Figures 2A, 2B and 2C, respectively.

It is known that the introduction of Ca, which is an acceptor-type dopant for LnNiO, leads to the formation of oxygen vacancies in the perovskite layer. As-formed oxygen vacancies are assumed to be occupied by the nearest interstitial oxygen, as the probability of the coexistence of both in the LnNiO structure is very low [67]. Thus, the overstoichiometric oxygen content in the samples decreases. The total oxygen content values for the LnCNCO samples are listed in Table S1. Despite the lower Ca content for the La-based series (15 mol.% for LCNO and 20 mol. % for PCNO and NCNO), the LCNO sample has the lowest oxygen content, which agrees well with the absolute oxygen content for undoped LnNiO, which was found to be 4.26, 4.24 and

4.16 for NNO, PNO and LNO, respectively [29]. Interestingly, for the PCNO samples with the same dopant level but synthesized via different techniques, the values of oxygen overstoichiometry are remarkably different. For example, for PCNO obtained via nitrate combustion with glycerol, similar to the results of this study, the delta is 0.12, whereas for PCNO obtained via hydroxide coprecipitation it is two times lower [68].

The TGA data presented in Figure 2 clearly demonstrate that the oxygen content in the samples decreases with increasing Cu doping. In the PCNCO series, the samples with $y \geq 0.2$ become oxygen deficient. The values obtained for Ca/Cu-co-doped PNO are lower than those obtained for Ca [68] and Cu-doped $\text{Pr}_2\text{NiO}_{4+\delta}$ [50] because both dopants tend to decrease the interstitial oxygen content.

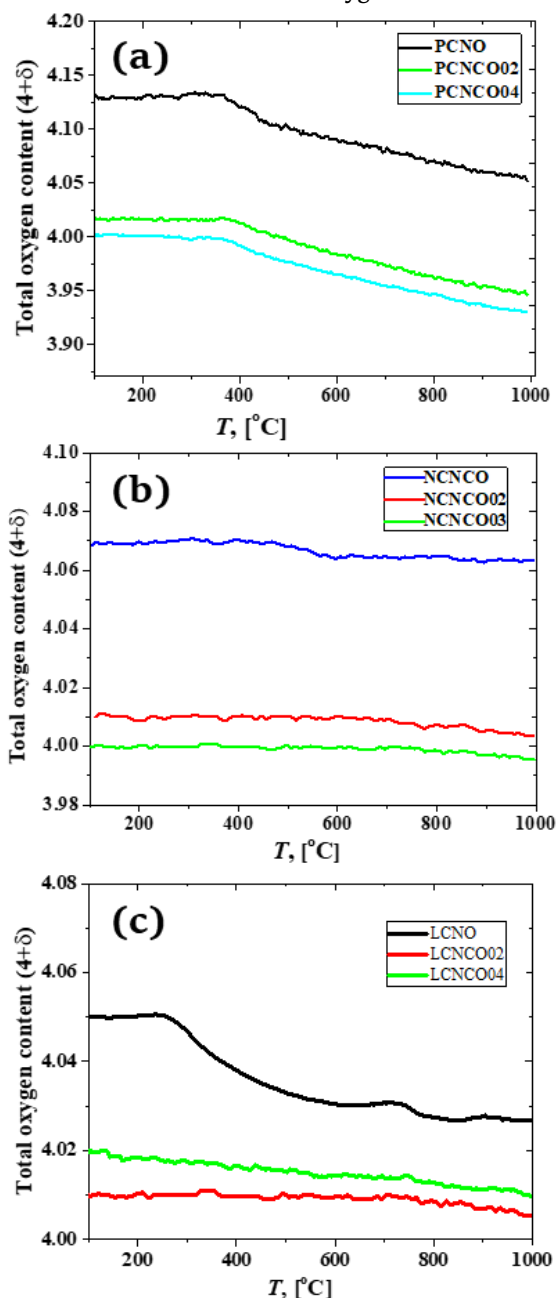


Figure 2 Total oxygen content in the materials of the PCNCO, NCNCO and LCNCO series, as calculated from the TGA data in air.

In the case of PCNCO, co-doping causes the oxygen content to decrease even faster than that in the NCNCO and LCNCO series, which remain overstoichiometric both at room and at the SOFC operating temperatures. The decrease in the absolute oxygen content can be explained by the fact that copper cations exhibit a lower average oxidation state in isostructural oxide compounds compared to nickel cations, as evident, for example, from the data on the oxygen nonstoichiometry of $\text{La}_2\text{CuO}_{4+\delta}$ [69] and $\text{La}_2\text{NiO}_{4+\delta}$ [70] under identical $p\text{O}_2$ - T conditions, as confirmed by the XPS studies of the PCNCO materials [57].

A further decrease with increasing temperature is related to the release of oxygen from the layered structure of the oxide. In Pr- and Nd-based systems, the release of oxygen results in the orthorhombic-to-tetragonal structure transition. The temperatures of this transition are listed in Table S1.

3.3. Conductivity

Conductivity was measured by a four-probe *dc* technique using ceramic samples of a bar shape. The relative densities of the samples, calculated as $\rho_{\text{exp}}/\rho_{\text{XRD}} \cdot 100\%$, were in the relatively narrow range of 92–94%, which was achieved by varying the sintering temperature as follows: 1450 °C for Cu-free samples, 1400 °C for $y = 0.1$, 1380 °C for $y = 0.2$, and 1350 °C for $y = 0.3$ – 0.4 . Figure 3 presents a summary of the total conductivity data for LnNiO and LnCNCO collected throughout the studies performed by our group [31, 40, 44, 56, 65, 68].

For convenience, the conductivity values at 700 °C are listed in Table S1.

As shown in Figure 3A, the temperature dependences of the total conductivity show semiconducting behavior at low temperatures, with a maximum at 350–500 °C, followed by a decrease in conductivity. The latter is mostly due to the beginning of active oxygen release from the studied samples, which is accompanied by a decrease in the concentration of electron holes, which are the major charge carriers in the oxides. The temperature of the maximum usually shifts to higher values with increasing doping.

For undoped LnNO, the highest conductivity value is observed for the system with the lowest lanthanide radius. It is known that the transfer of charge carriers in layered nickelates occurs mainly in the equatorial planes of the NiO_6 octahedra [1]. A decrease in the Ni-O1 bond length with a decrease in the radius of the Ln^{3+} ion results in an increase in the intensity of jumps; this is associated with an increase in electrical conductivity and the maximum electrical conductivity attained for the NNO sample.

Depending on the composition of the doped samples, the following dependences are observed: 1) doping with Ca^{2+} leads to an increase in electrical conductivity, and 2) co-doping with Cu^{2+} results in a decrease in conductivity.

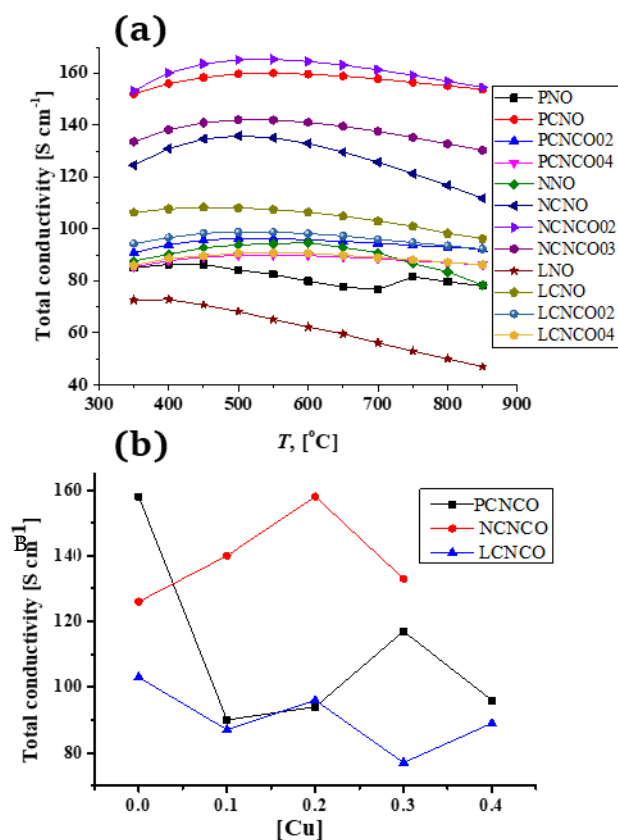


Figure 3 Total conductivity in air for the LnNO and LnCNCO series (Ln = Pr, Nd, La): temperature dependences (A) and concentration dependences at 700 °C.

The substitution of Ca²⁺ for Ln³⁺ cations lead to an increase in the concentration of compensating electron holes, which leads to an increase in the conductivity level compared with LnNO due to an increase in the electron component of the total conductivity. Co-doping with Cu results in a decrease in the conductivity of the LaCNCO series (Figure 3B), which correlates well with the literature data [49]. However, for the Nd/Ca and Pr/Ca series (except for PCNO), such a decrease can be observed only for the sample with 40 mol.% copper ($y = 0.4$), which can be explained by several factors: a decrease in the conductivity due to a decrease in the concentration of the electronic charge carriers-electron holes localized on the B-site transition metal cations (i.e., the concentration of Ni³⁺) due to Cu²⁺ doping and improved sinterability of the Cu-containing materials. The conduction behavior of Nd- and Pr-based systems may also be affected by the phase transition and the miscibility gap [71, 72].

3.4. Oxygen mobility

The TPIE C¹⁸O₂ curves are given in Figure 4. The isotope exchange process starts at ~250 °C and involves ~90–100% of the total oxygen content of samples, thereby exhibiting favorable oxygen transport characteristics. The $\alpha(T)$ curves for all the samples possess a few extrema, indicating the presence of multiple oxygen forms with varying isotope substitution rates within the samples.

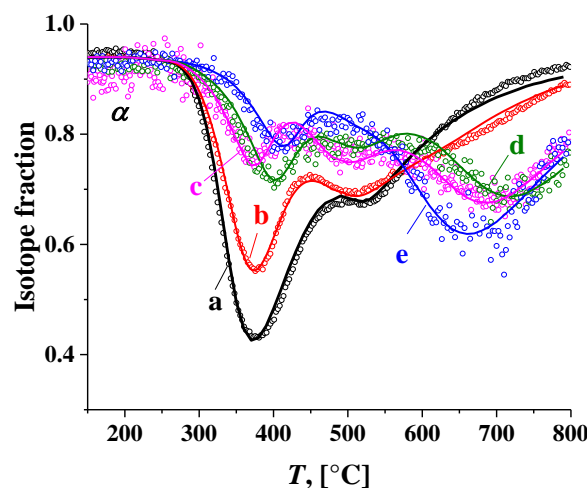


Figure 4 Temperature-programmed isotope exchange of oxygen with C¹⁸O₂ in the flow reactor for the Pr_{1.6}Ca_{0.4}Ni_{1-y}Cu_yO_{4+δ} samples, $y = 0.0$ (a), 0.1 (b), 0.2 (c), 0.3 (d) and 0.4 (e). Points – experiment, lines – modeling.

In the case of the Pr_{1.6}Ca_{0.4}NiO_{4+δ} sample, two clearly distinguishable extrema are evident at temperatures of approximately 350 and 540 °C. These extremes correspond to two distinct oxygen forms, which can be designated O_{fast} and O_{middle}, respectively. For the Pr_{1.6}Ca_{0.4}Ni_{0.9}Cu_{0.1}O_{4+δ} sample, two extrema are observed at almost the same temperatures. However, the area of the first extremum is found to be lower compared to that of the Pr_{1.6}Ca_{0.4}NiO_{4+δ} sample, indicating that the dynamics of reaching isotope equilibrium are slower. The doping results in a reduction in the amount of fast oxygen, accompanied by the emergence of very slowly substituting forms of oxygen (O_{slow}) (at $y \geq 0.2$). As the Cu-content increases, the amount of fast oxygen decreases, leading to a decrease in the overall rate of oxygen substitution.

A mathematical model considering the mechanism of isotope exchange, in which there is a single diffusion channel from the surface into the oxide bulk involving the most mobile oxygen and then subsequent isotope exchange occurs with neighboring, more strongly bound lattice oxygen species, was successfully used for Ca- and Cu-doped lanthanum nickelates [42] and many other complex oxides [12, 13]. However, its application is challenging with respect to the precise representation of the experimental TPIE C¹⁸O₂ data obtained for the Pr_{1.6}Cu_{0.4}Ni_{1-y}Cu_yO_{4+δ} series. A more accurate description can be achieved via a mathematical model that includes two or three distinct oxygen forms, each characterized by a specific diffusion coefficient value (Figure 5). The estimated fractions of each oxygen form, along with the corresponding diffusion coefficients and activation energy values, are provided in Table 3.

For the Pr_{1.6}Ca_{0.4}NiO_{4+δ} and Pr_{1.6}Ca_{0.4}Ni_{0.9}Cu_{0.1}O_{4+δ} samples, surface exchange was identified as the limiting stage at temperatures below 300 °C. The surface exchange rate can be described by a mass exchange (or interphase exchange) coefficient K (Figure 5) since the precise limiting factor is unclear due to the small specific surface area of the samples (Table 1).

Table 3 Parameters of isotope exchange for $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ ($y = 0.0-0.4$) at 423 °C: mass exchange coefficient (K), fractions of different oxygen forms (θ), tracer diffusion coefficient (D^*), and effective activation energy (E_a).

y	Surface			Volume							
	$K \cdot 10^4$, [mol/(m ² s)]	E_a , [kJ/mol]	θ , [%]	O_{fast}			O_{middle}			O_{slow}	
				$D^* \cdot 10^{11}$, [cm ² /s]	E_a , [kJ/mol]	θ , [%]	$D^* \cdot 10^{13}$, [cm ² /s]	E_a , [kJ/mol]	θ , [%]	$D^* \cdot 10^{15}$, [cm ² /s]	E_a , [kJ/mol]
0.0	0.8	160	77	3.6		23	0.6	160	0	-	-
0.1	22	160	40	2.3		30	0.5	160	30 ^a	(0.1÷1)	160
0.2	Not limited		15	4.2	160	25 ^a	(0.2÷2)	160	60	1.0	160
0.3	Not limited		20	1.7		20 ^a	(0.1÷1)	160	60	0.7	160
0.4	Not limited		13	0.7		7 ^a	(0.1÷1)	160	80	1.1	160

^a These forms of oxygen are actually a set of the oxygen forms close by their mobility. In this case, a range of diffusion coefficient values is estimated.

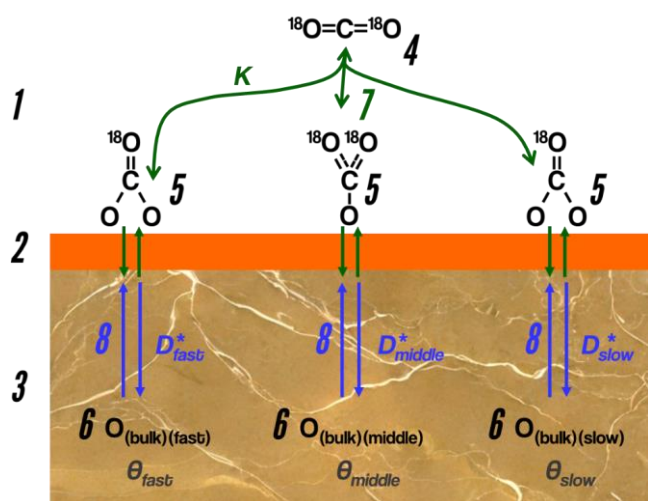


Figure 5 Schematic illustration of the model of oxygen transport in $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$. 1 - gas phase (including pores/voids in the sample), 2 - sample surface, 3 - sample bulk, 4 - C^{18}O_2 molecule, 5 - surface carbonate complexes, 6 - bulk oxygen forms with fast, middle and slow transport and fractions of θ_{fast} , θ_{middle} and θ_{slow} , respectively, 7 - mass exchange between gas and solid phases (including oxygen surface exchange and permeation of the surface layer) characterized by a coefficient K , 8 - fast, middle and slow diffusion channels characterized by oxygen tracer diffusion coefficients D_{fast} , D_{middle} and D_{slow} , respectively.

This may be due to the rate of exchange with surface oxygen species, the permeability of the surface layer, or a combination of these factors.

The Arrhenius plots for the oxygen mass exchange (K) and tracer diffusion coefficients (D^*) are given in Figures 6 and 7, respectively. The higher K values observed for the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.9}\text{Cu}_{0.1}\text{O}_{4+\delta}$ sample as compared with those for the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{NiO}_{4+\delta}$ one is probably related to variations in the surface heteroexchange rate or permeability of the sample surface layer, which can be caused by the sample surface properties, such as the effective charge of the cations and the contributions of more strongly bound surface oxygen [43], as well as the effects of the sample dispersion and the porous structure [45, 46].

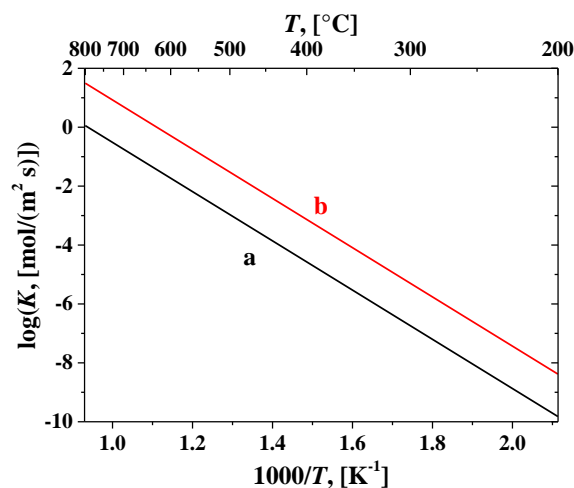


Figure 6 Arrhenius plots for the mass exchange coefficient acquired for the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ samples from TPIE C^{18}O_2 data, $y = 0.0$ (a) and 0.1 (b).

The oxygen tracer diffusion coefficient values and the contribution of fast oxygen forms tend to decrease with increasing Cu content, which is similar to the trends observed and discussed for Ca- and Cu-doped La [44] and Nd nickelates [47] and can be explained by the variation in highly mobile interstitial oxygen content, since the effective activation energy values remain the same within the calculation error. However, the highest D^* value for O_{fast} is demonstrated for the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.8}\text{Cu}_{0.2}\text{O}_{4+\delta}$ sample ($9.4 \cdot 10^{-8}$ cm²/s at 700 °C (Figure 7, dependence d), which is slightly higher than that for the Cu-free sample (Figure 7, dependence b).

This maximum D^* value for O_{fast} correlates well with the minimum polarization resistance observed for the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.8}\text{Cu}_{0.2}\text{O}_{4+\delta}$ electrode with a collector ($0.27 \Omega \text{ cm}^2$ at 700 °C) in our recent study, which is lower than the values obtained for PCNO ($0.64 \Omega \text{ cm}^2$ [57], $0.92 \Omega \text{ cm}^2$ [68]) and PNO ($0.72 \Omega \text{ cm}^2$ [68]). On the other hand, the contribution of O_{fast} is lower than that of the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.7}\text{Cu}_{0.3}\text{O}_{4+\delta}$ sample (Table 2).

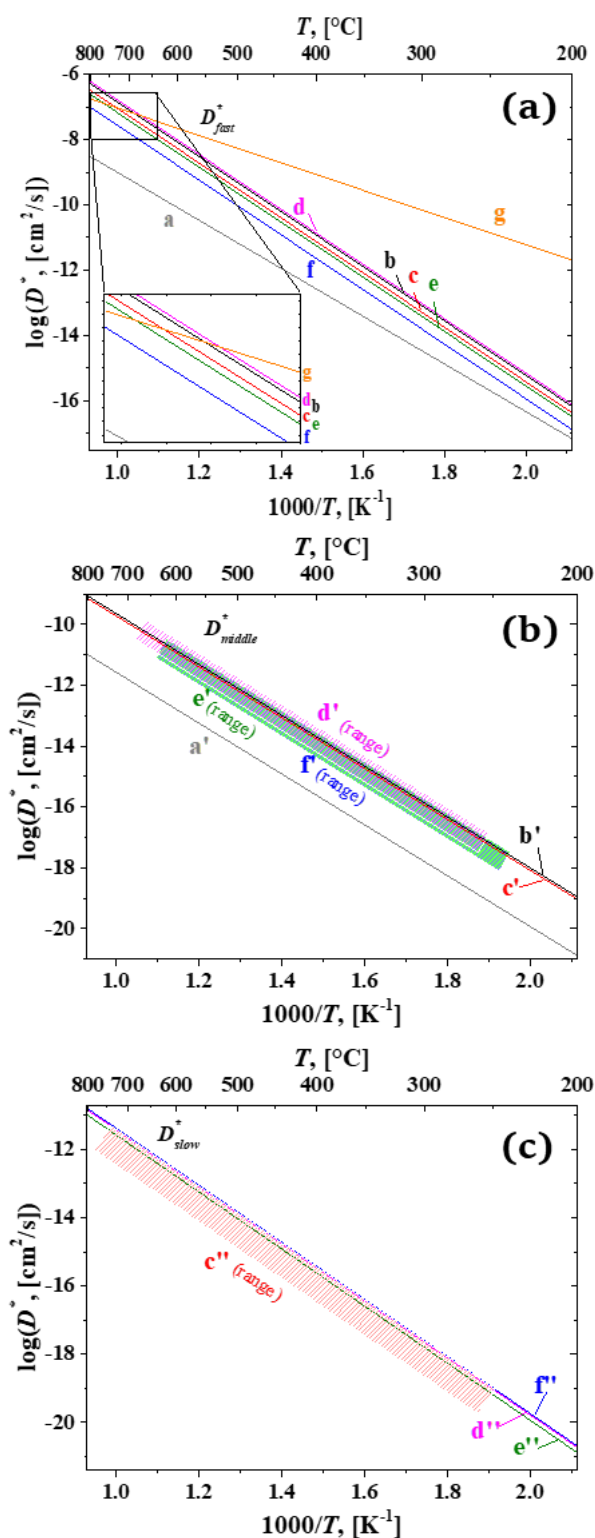


Figure 7 Arrhenius plots for the oxygen tracer diffusion coefficient (Fast (A), Middle (B) and Low (C)) acquired for the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{1-y}\text{Cu}_{0.4}\text{O}_{4+\delta}$ samples from the TPIE C^{18}O_2 data, $y = 0.0$ (a, a') [68], 0.0 (b, b'), 0.1 (c, c'), 0.2 (d, d'), 0.3 (e, e') and 0.4 (f, f'), compared with the undoped $\text{Pr}_2\text{NiO}_{4+\delta}$ (g) [68].

Nevertheless, the related electrode demonstrates worse electrochemical properties in the series of Cu-doped PCNO [57], which may be related to the nonoptimized electrode microstructure. From our point of view, the potential of this material should be further explored, as it was performed for $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.6}\text{Cu}_{0.4}\text{O}_{4+\delta}$ [58].

Notably, the diffusion coefficient values for the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{NiO}_{4+\delta}$ sample acquired in this work are ~ 2 orders of magnitude greater than those of the sample with the same composition studied earlier [68], which can be associated with differences in synthesis technique and sintering temperature. These preparation parameters affect not only the dispersity of the powders but also the interstitial oxygen content, which is a crucial factor influencing the oxygen kinetic features. The δ values obtained for the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{NiO}_{4+\delta}$ sample synthesized via coprecipitation of hydroxides [68] and nitrate combustion in this study are 0.06 and 0.12, respectively. As expected, the higher interstitial oxygen content results in superior diffusion properties (Table 4).

The oxygen tracer diffusion coefficient values corresponding to O_{fast} are lower than those of undoped $\text{Pr}_2\text{NiO}_{4+\delta}$ (Figure 7, dependence g) at low and intermediate temperatures; however, they are comparable or even greater at high temperatures. Nevertheless, the oxygen mobility remains high for all samples in the PCNCO series ($D_{\text{fast}} \sim 10^{-8} \text{ cm}^2/\text{s}$ at 700°C), which makes them promising materials for practical application. For comparison, the data for Cu-doped LNCNO are summarized in Table 4 and compared with the data obtained for undoped oxides. It is seen that in the NCNCO series, a weak dependence of the diffusion properties on the Cu content was observed; all the samples demonstrate D_{fast} values of $\sim 10^{-9} \text{ cm}^2/\text{s}$ at 700°C , whereas for the LNCNO series, the D_{fast} is one order of magnitude lower.

It is well known that the oxygen transport in Ruddlesden–Popper phases proceeds via a cooperative mechanism involving both regular and highly mobile interstitial oxygen through interstitialcy movement [26, 30, 75, 76]. Doping A-site with larger cations such as Ca leads to the hampering of the cooperative mechanism of oxygen migration due to the effect of steric hindrances for an oxygen jump and a decrease in the interstitial oxygen content, leading to decreasing D^* values and sometimes the emergence of slow diffusion channels, which is discussed in detail in our previous works [14, 29, 35, 68, 77]. Therefore, the PCNO samples, both of which were studied earlier [28] (Figure 7, dependence a) and in this work (Figure 7, dependence b), have two oxygen forms characterized by two different diffusion coefficients. O_{fast} characterized by D_{fast} apparently corresponds to the motion of oxygen via the cooperative mechanism. Less mobile oxygen form O_{middle} probably corresponds to the oxygen moving via the hampered cooperative mechanism [35]. Alternatively, O_{slow} can be ascribed to oxygen diffusion via the vacancy mechanism in the perovskite layers, which become oxygen deficient at room temperature (with δ values equal to -0.03 , -0.03 and 0.00 at $y = 0.2$, 0.3 and 0.4 , respectively), and further vacancy formation occurs with increasing temperature. Indeed, in the works [78–81], the formation of oxygen vacancies for first-order RP phases with low, close-to-zero or negative oxygen overstoichiometry δ was demonstrated.

Table 4 Parameters of isotope exchange for $\text{Ln}_{2-x}\text{Ca}_x\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ at 700 °C: tracer diffusion coefficient (D^*) and effective activation energy (E_a) [44, 47, 68, 73, 74].

Sample	Volume						Ref.
	O_{fast}		O_{middle}		O_{slow}		
	D^* , [cm ² /s]	E_a , [kJ/mol]	D^* , [cm ² /s]	E_a , [kJ/mol]	D^* , [cm ² /s]	E_a , [kJ/mol]	
LNO	$2.0 \cdot 10^{-9}$	100	-	-	-	-	[73]
LCNO	$6.5 \cdot 10^{-10}$						
LCNO	$1.4 \cdot 10^{-9}$						
LCNCO01	$1.1 \cdot 10^{-9}$						
LCNCO02	$7.5 \cdot 10^{-10}$	100	-	-	-	-	[44]
LCNCO03	$6.2 \cdot 10^{-10}$						
LCNCO04	$5.0 \cdot 10^{-10}$						
PNO	$6.9 \cdot 10^{-8}$	80	-	-	-	-	[68]
PCNO	$6.1 \cdot 10^{-10}$	140	$1.6 \cdot 10^{-12}$	160	-	-	
PCNO	$8.0 \cdot 10^{-8}$		$1.2 \cdot 10^{-10}$		-	-	
PCNCO01	$5.2 \cdot 10^{-8}$		$1.1 \cdot 10^{-10}$		$\sim 10^{-13}$		
PCNCO02	$9.4 \cdot 10^{-8}$	160	$\sim 10^{-10}$	160	$2.2 \cdot 10^{-12}$	160	This work
PCNCO03	$3.8 \cdot 10^{-8}$		$\sim 10^{-10}$		$1.6 \cdot 10^{-12}$		
PCNCO04	$1.6 \cdot 10^{-8}$		$\sim 10^{-10}$		$2.5 \cdot 10^{-12}$		
NNO	$4.5 \cdot 10^{-8}$	140	-	-	-	-	[74]
NCNO	$9.8 \cdot 10^{-10}$	120					
NCNO							
NCNCO							
NCNCO02	$3.3 \cdot 10^{-9}$	120	-	-	-	-	[47]
NCNCO03							

Additionally, the O_{middle} and O_{slow} oxygen forms can be associated with the effect of local distortion of the perovskite structure in the vicinity of the Cu cations, which, in turn, can affect the binding energy of oxygen [42]. Notably, despite the higher δ value for the first member of the series (PCNO) determined in this study (0.12), the introduction of Cu cations resulted in a more significant decrease in the interstitial oxygen content compared to the NCNCO and LCNCO series of co-doped nickelates, which remain overstoichiometric at room temperature up to $y = 0.3$ and 0.4 , respectively [40, 56]. Nevertheless, the appearance of oxygen vacancies in the perovskite structure of layered oxides can be favorable for electrode performance, as shown in [40].

The oxide ionic conductivity (σ_i) of the PCNCO series is estimated from D^* values via the Nernst–Einstein equation (Equation (4)):

$$\sigma_i = \frac{4q_e^2 N_0}{f k_B T} D^*, \quad (4)$$

where f is the correlation factor (taken to be equal to 1); q_e is the electronic charge; N_0 is the total concentration of the oxygen ions in the substance; and k_B is the Boltzmann constant.

The σ_i values tend to decrease with increasing Cu content due to decreasing D_{fast} values and O_{fast} contributions (Figure 8A). Nevertheless, the values ($\sim 10^{-3}$ – 10^{-2} S/cm at

700 °C) are close to those for other RP phases [74, 82]. It should be also noted that for the PCNCO series the ionic conductivity values are superior to those for NCNCO and LCNCO. For example, for samples with a medium copper content ($y = 0.2$), the ionic conductivity values at 700 °C were evaluated to be $3.7 \cdot 10^{-3}$, $4.08 \cdot 10^{-4}$ and $1.05 \cdot 10^{-4}$ S/cm for $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.8}\text{Cu}_{0.2}\text{O}_{4+\delta}$, $\text{Nd}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.8}\text{Cu}_{0.2}\text{O}_{4+\delta}$ and $\text{La}_{1.7}\text{Ca}_{0.3}\text{Ni}_{0.8}\text{Cu}_{0.2}\text{O}_{4+\delta}$, respectively.

When the oxide ionic conductivity is compared with the total conductivity [57], one can conclude that electronic transport predominates (O^{2-} transfer number $\sim 10^{-3}$ at 700 °C); hence, the ambipolar conductivity, $\sigma_{\text{amb}} = \frac{\sigma_e \sigma_i}{(\sigma_e + \sigma_i)}$, can be estimated to be almost equal to the oxide ionic conductivity. Hence, the ambipolar conductivity value series can be considered to be high enough for practical application.

3.5. Electrochemical properties

Table 5 summarizes the electrochemical properties of LnCNCO-based electrodes in contact with $\text{Ce}_{0.2}\text{Sm}_{0.2}\text{O}_{1.9}$ electrolyte, which were recently studied elsewhere [40, 47, 57]. It can be seen, that the electrodes with PCNCO functional layers demonstrate the most promising characteristics, which was supported by the enhanced oxygen kinetic properties obtained in this study. On the contrary, a correlation with the electrical properties of the materials is not apparent. This is a distinctive feature of the MIEC materials [59].

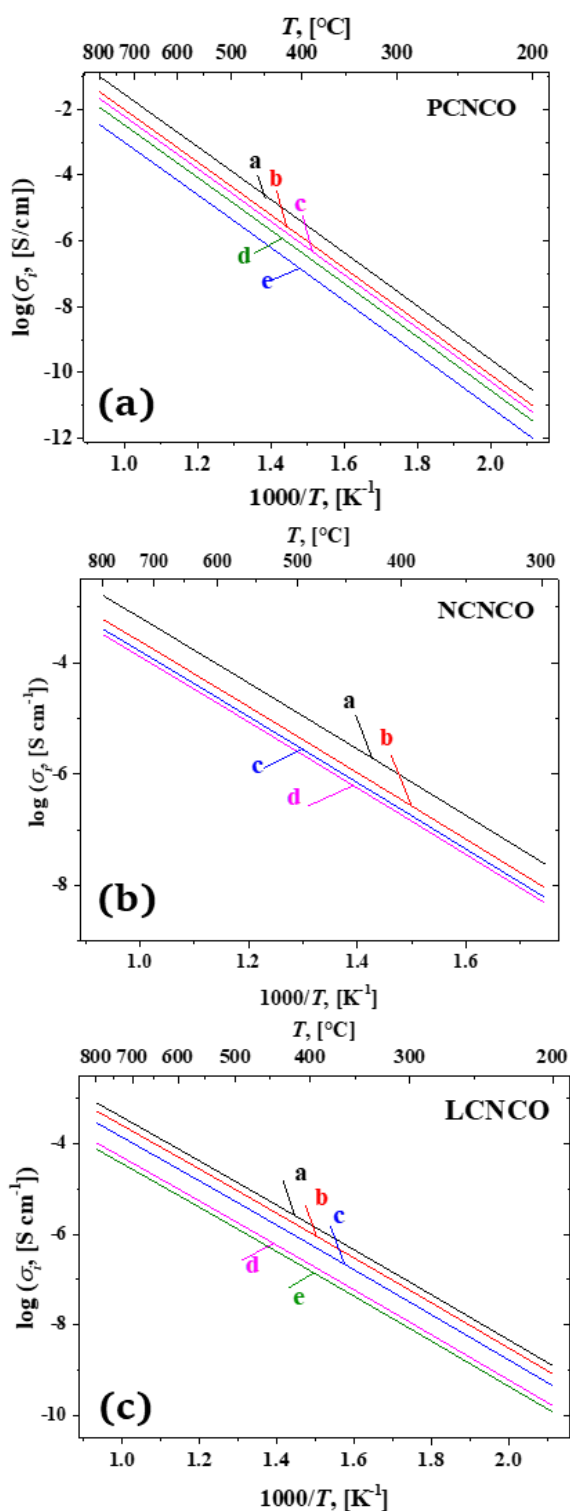


Figure 8 Oxide ionic conductivity estimated from the isotope exchange data: A) PCNCO, $y = 0.0$ (a), 0.1 (b), 0.2 (c), 0.3 (d) and 0.4 (e); B) NCNCO, $y = 0.0$ (a), 0.1 (b), 0.2 (c), 0.3 (d) and C) LCNCO, $y = 0.0$ (a), 0.1 (b), 0.2 (c), 0.3 (d) and 0.4 (e).

The analysis of the data presented in Table 5 shows that the electrodes with the PCNCO functional layers exhibited the lowest polarization resistance at $y = 0.1$ and 0.2 (1.31 and $1.36 \Omega \text{ cm}^2$ at $700 \text{ }^\circ\text{C}$, respectively [57]). In contrast, the polarization resistance of the LCNCO electrodes demonstrated an increase at low Cu content ($y \leq 0.2$), followed by a subsequent decrease. Therefore, the LCNCO4 electrode

exhibited the optimal performance in the LCNCO series ($2.01 \Omega \text{ cm}^2$ at $700 \text{ }^\circ\text{C}$, [40]). The lowest value in the NCNCO series was observed at $y = 0.2$ ($5.70 \Omega \text{ cm}^2$ at $700 \text{ }^\circ\text{C}$, [47]). The application of a highly conductive oxide collector based on $\text{LaNi}_{0.6}\text{Fe}_{0.4}\text{O}_{3-\delta}$ was found to increase the electrode performance due to better current distribution across the electrode volume. As, for example, for the PCNCO2 electrode, the application of the collector allowed decreasing R_p by approximately four times, down to $0.27 \Omega \text{ cm}^2$ at $700 \text{ }^\circ\text{C}$.

Moreover, as we have demonstrated recently [57], the PCNCO electrodes with LNF collector may be successfully used in contact with proton-conducting electrolytes, exhibiting the R_p values in the range of 0.19 – $0.21 \Omega \text{ cm}^2$ at $700 \text{ }^\circ\text{C}$ [57].

4. Limitations

The assumptions on the oxygen transport mechanisms for each oxygen form for the studied materials are based on the experimental data as well as the data presented in the literature, including the authors' previous works. However, elucidating the nature of these oxygen forms and details of their migration pathways requires further studies, which can be carried out in the future.

5. Conclusions

This work continues a series of investigations concerning the physicochemical properties, oxygen diffusion and electrochemical performance of $\text{Ln}_{2-x}\text{Ca}_x\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ ($\text{Ln} = \text{Pr}, \text{Nd}, \text{La}$) materials. In particular, this study aimed to investigate the oxygen content and diffusion in $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ ($y = 0.0$ – 0.4) oxides obtained via nitrate combustion with glycerol as a fuel. The materials were found to exhibit an orthorhombic structure, related to the absolute oxygen content in the range of 4.15 – 3.97 . The oxygen mobility in the samples was investigated via temperature-programmed isotope exchange of oxygen (TPIE) with C^{18}O_2 in a flow reactor. The obtained curves, comprising few extrema, were fitted via a mathematical model that included two or three distinct oxygen forms, each characterized by a specific diffusion coefficient value: (1) O_{fast} , characterized by D_{fast} corresponding to the oxygen moving via the cooperative mechanism, (2) O_{middle} , corresponding to the oxygen moving via a hampered cooperative mechanism, and (3) O_{slow} , ascribed to the oxygen diffusion via a vacancy mechanism in the perovskite layers, which became oxygen deficient at IT-SOFC operating temperatures. The oxygen tracer diffusion coefficient values and the contribution of fast oxygen forms tend to decrease with increasing Cu content, which is similar to the trends observed and discussed for Ca- and Cu-co-doped La and Nd nickelates in our recent studies. These features can be explained by the variation in the highly mobile interstitial oxygen content, since the effective activation energy values remained approximately the same within the calculation error.

Table 5 Polarization resistance of the electrodes with $\text{Ln}_{2-x}\text{Ca}_x\text{Ni}_{1-y}\text{Cu}_y\text{O}_{4+\delta}$ functional layers measured in symmetrical cells with the $\text{Ce}_{0.8}\text{Sm}_{0.2}\text{O}_{1.9}$ electrolyte (at 700 °C in air). Electrodes without and with a collector layer based on $\text{LaNi}_{0.4}\text{Fe}_{0.6}\text{O}_{3-\delta}$ (LNF) are presented [40, 47, 57, 58, 73, 74, 77, 83].

Electrode FL	without collector		with LNF collector		Ref.
	R_p , [$\Omega \text{ cm}^2$]	E_a , [kJ/mol]	R_p , [$\Omega \text{ cm}^2$]	E_a , [kJ/mol]	
$\text{La}_2\text{NiO}_{4+\delta}$			0.57	105	
$\text{La}_{1.7}\text{Ca}_{0.3}\text{NiO}_{4+\delta}$			4.01	102	[73]
$\text{La}_{1.7}\text{Ca}_{0.3}\text{NiO}_{4+\delta}$	10.52	162			
$\text{La}_{1.7}\text{Ca}_{0.3}\text{Ni}_{0.9}\text{Cu}_{0.1}\text{O}_{4+\delta}$	14.36	139			
$\text{La}_{1.7}\text{Ca}_{0.3}\text{Ni}_{0.8}\text{Cu}_{0.2}\text{O}_{4+\delta}$	12.76	150			[40]
$\text{La}_{1.7}\text{Ca}_{0.3}\text{Ni}_{0.7}\text{Cu}_{0.3}\text{O}_{4+\delta}$	6.87	138			
$\text{La}_{1.7}\text{Ca}_{0.3}\text{Ni}_{0.6}\text{Cu}_{0.4}\text{O}_{4+\delta}$	2.01	133	2.65	140	
$\text{Pr}_2\text{NiO}_{4+\delta}$			0.72		
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{NiO}_{4+\delta}$			0.93		[77]
$\text{Pr}_2\text{NiO}_{4+\delta}$			0.42	112	
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{NiO}_{4+\delta}$			0.74	141	[83]
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{NiO}_{4+\delta}$	1.48	129	0.64	118	
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.9}\text{Cu}_{0.1}\text{O}_{4+\delta}$	1.31	131			
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.8}\text{Cu}_{0.2}\text{O}_{4+\delta}$	1.36	121	0.27	117	[57]
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.7}\text{Cu}_{0.3}\text{O}_{4+\delta}$	1.65	116			
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.6}\text{Cu}_{0.4}\text{O}_{4+\delta}$	1.48	92	0.37	77	
$\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.6}\text{Cu}_{0.4}\text{O}_{4+\delta}$	0.60	86	0.38	99	[58]
$\text{Nd}_2\text{NiO}_{4+\delta}$			0.71	116	
$\text{Nd}_{1.6}\text{Ca}_{0.4}\text{NiO}_{4+\delta}$			0.38	111	[74]
$\text{Nd}_{1.6}\text{Ca}_{0.4}\text{NiO}_{4+\delta}$	12.28	98	4.21		
$\text{Nd}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.9}\text{Cu}_{0.1}\text{O}_{4+\delta}$	10.55	117	3.81		
$\text{Nd}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.8}\text{Cu}_{0.2}\text{O}_{4+\delta}$	5.70	102	3.32		[47]
$\text{Nd}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.7}\text{Cu}_{0.3}\text{O}_{4+\delta}$	15.69	151	6.91		

The highest D^* value for O_{fast} was demonstrated for the $\text{Pr}_{1.6}\text{Ca}_{0.4}\text{Ni}_{0.8}\text{Cu}_{0.2}\text{O}_{4+\delta}$ sample, which was $9.4 \cdot 10^{-8} \text{ cm}^2/\text{s}$ at 700 °C. Co-doping with Cu results in a decrease in the conductivity of the LaCNCO series; however, for the Pr/Ca and Nd/Ca series, such a decrease was observed only for the samples with 40 mol.% copper ($y=0.4$), which can be explained by several factors, such as a decrease in the conductivity due to the reduction in the concentration of the electronic charge carriers and electron holes localized on the B-site transition metal cations (i.e., the concentration of Ni^{3+}) due to Cu^{2+} doping and improved sinterability of the samples from the Cu-containing materials. Despite the decrease in conductivity, an increase in the electrochemical performance was demonstrated for the Pr-based electrodes with moderate calcium and copper contents ($x=0.4$, $y=0.2$). This enhancement was observed for the La-based series only at $y=0.4$, whereas the performance of the Nd-based electrodes co-doped with Ca and Cu was inferior to that of the undoped material, which can be explained by the non-optimized electrode microstructure.

• Supplementary materials

This manuscript contains supplementary materials, which are available on the corresponding online page.

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Resources: T.Z.

Supervision: V.S.

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● Conflict of interest

The authors declare that they have no conflicts of interest.

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RAS, <https://ihte.ru/>;

Ural Federal University, <https://urfu.ru/en/>;

Institute of Metallurgy UB RAS, <http://www.imet-uran.ru/>.

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