


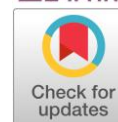


Composite catalysts based on the CaO–Fe₂O₃ system for the oxidative conversion of methane

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Abstract

The catalysts of the CaO–Fe₂O₃ system with Fe₂O₃ content in the range of 0–100 wt.% were synthesized by the solid state method at 900 and 1000 °C. The catalysts were characterized by XRD and SEM–EDX methods, and their activity in the oxidative conversion of methane at 750 °C was studied. The phase composition of the catalysts corresponds to the CaO–Ca₂Fe₂O₅, Ca₂Fe₂O₅–CaFe₂O₄ and CaFe₂O₄–α-Fe₂O₃ regions of the phase diagram. The specific catalytic activity dependence on the Fe₂O₃ content has an extremum. The CaO–Ca₂Fe₂O₅ catalysts have the highest reactivity and the active centers in them are localized at the phase interface. The activity of Ca₂Fe₂O₅–CaFe₂O₄ samples decreases with an increase in the CaFe₂O₄ content. The CaFe₂O₄–α-Fe₂O₃ catalysts have a core-shell structure and exhibit the least activity, which is determined by the CaFe₂O₄ shell.

Keywords

catalysts
calcium ferrites
solid state synthesis
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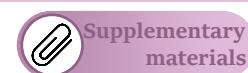
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Key findings

- Of the studied compositions in the CaO–Fe₂O₃ system, the catalysts CaO–Ca₂Fe₂O₅ have the greatest activity in the methane oxidative conversion.

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1. Introduction

Oxide systems based on the calcium ferrites due to such physicochemical properties as high-temperature stability, ability to transport oxygen ions, and mixed ion-electron conductivity are widely studied as catalysts for the hydrocarbons and CO oxidation [1, 2] and as oxygen carriers for high-temperature cyclic processes, including fuels combustion [3], gasification of coal [4–6] and biomass [7, 8], hydrogen generation [9, 10], etc. An important advantage of calcium ferrite based catalysts over other oxidation catalysts is their high activity at low cost. However, systematic studies of the relationship between the composition, structural characteristics and activity of calcium ferrite catalysts have not been carried out. In this paper, the influence of the composition of the CaO–Fe₂O₃ catalysts with varying Fe₂O₃ content in the range of 0–100 wt.% on the activity in the oxidative methane conversion was studied.

2. Experimental

The catalysts of the CaO–Fe₂O₃ system were synthesized using Fe₂O₃ of the "high purity" qualification (TU 6-09-1418-78) and CaO of the "pure" qualification (GOST 8677-66). The reagents were mixed and ground in the required proportion using a KM-1 ball mill for an hour; the ground sample was pressed (pressure of 3.5·10³ kg/cm², exposure time of 1.5–3 min) into tablets with a diameter of 17 mm and a thickness of 1–2 mm. The tablets were calcined at 900 °C for 10 h or at 1000 °C for 4 h, and then cooled to room temperature at a rate of ~8 °C/min. The resulting tablets were crushed to obtain a fraction of 0.1–0.2 mm. The samples were denoted as xx% Fe₂O₃, where xx is the mass content of Fe₂O₃ in the initial charge, for example, a sample of 57% Fe₂O₃ contains 57 wt.% Fe₂O₃, and a sample of 100% Fe₂O₃ corresponds to pure hematite treated according to the described procedure.

The quantitative phase composition of the samples and the microstructural characteristics of the phases were determined using an X'Pert PRO MPD diffractometer (PANalytical, Netherlands) with a PIXcel solid-state detector (Co K α) at room temperature and 750 °C. The results were processed using a full-profile analysis of polycrystalline substances using the Rietveld method and the method of minimizing the derivative difference [11].

The specific surface area of the samples was determined using the NOVA 3200e sorption analyzer (Quantachrome Instruments, USA) in the low-temperature nitrogen adsorption mode at -195.8 °C. The calculation of the specific surface area ($S_{sp.}$) was performed using the modified BET method in accordance with the international standard [12].

The surface composition and morphology of the catalyst particles were studied using a TM-3000 and TM-4000 scanning electron microscopes (Hitachi, Japan) with an energy dispersive X-ray spectrometer (EDX) Bruker XFlash 430H (Germany) equipped with a Quantax 70 microanalysis system. The accumulation time, which was at least 10 min, was determined by the quality of the spectrum assembly, which allows quantitative processing. The polished tablets or fractions of 0.1–0.2 mm of the samples were examined.

The catalytic activity in the oxidative methane conversion was studied using the continuous-flow fixed-bed unit with a stationary catalyst layer in a quartz microreactor, with of a mixture of CH₄:O₂:He = 82:9:9vol.%, at a temperature of 750±0.5 °C. The analysis of the initial gas mixture and reaction products was carried out using an Agilent 7890A GC chromatograph (Agilent Technologies Inc., USA) with the following columns: "Molecular Sieve 5A", "HP Plot Al₂O₃", "DB-1" and "Haysep Q". The specific rates of CO₂, CO and C₂-hydrocarbons formation were determined at the degree of oxygen conversion $X_{O_2} = 5-10\%$.

3. Results and Discussion

According to phase diagram of the CaO-Fe₂O₃ system, the ranges 0–58.7, 58.7–74 and 74–100 wt.% of Fe₂O₃ up to 1155 °C correspond to the regions of the phase compositions CaO-Ca₂Fe₂O₅, Ca₂Fe₂O₅-CaFe₂O₄ and CaFe₂O₄- α -Fe₂O₃. There are stable calcium ferrites of stoichiometric composition – dicalcium ferrite, Ca₂Fe₂O₅, containing 58.7 wt.% Fe₂O₃, and monocalcium ferrite, CaFe₂O₄ containing 74 wt.% Fe₂O₃.

The results of the XRD phase analysis of the catalysts ($T_{synt.} = 900$ and 1000 °C) showed a good correspondence of the Ca₂Fe₂O₅-CaFe₂O₄ and CaFe₂O₄- α -Fe₂O₃ samples to the phase diagram and a slight difference in the compositions of CaO-Ca₂Fe₂O₅ catalysts from those seen in the phase diagram. The presence of CaFe₂O₄ phase in the CaO-Ca₂Fe₂O₅ samples, $T_{synt.} = 900$ °C, (Figure 1) is due to kinetic limitations, since the Ca₂Fe₂O₅ ferrite formation proceeds through the formation of CaFe₂O₄ ferrite [13].

This is confirmed by the fact that among the samples in synthesized at 1000 °C, CaFe₂O₄ ferrite was detected only when the Fe₂O₃ content was greater than 45 wt.%. The XRD patterns of 90, 65 and 45% Fe₂O₃ samples ($T_{synt.} = 1000$ °C) are shown in Figures S1–S3.

The specific surface area ($S_{sp.}$) values of the samples synthesized at 900 °C vary between 0.8 and 1.5 m²/g. At $T_{synt.} = 1000$ °C samples have a slightly smaller $S_{sp.}$ (0.2–1.0 m²/g), while the surface of the samples with the composition CaO-Ca₂Fe₂O₅ varies little, within the ranges of 1.1–1.5 (900 °C) and 0.8–1.0 (1000 °C) m²/g, respectively.

In accordance with the SEM images, the morphology of the synthesized catalysts differs. Figure 2 shows the change in sample images with an increase in the content of Fe₂O₃ in CaO-Fe₂O₃. Taking into account the phase composition of samples and the results of EDX analysis, the following conclusions were drawn. Practically single-phase samples 57% Fe₂O₃ (93% Ca₂Fe₂O₅) and 74% Fe₂O₃ (98% CaFe₂O₄) have different morphological structures (Figure 2). The 57% Fe₂O₃ sample has an almost homogeneous structure, consists of sintered particles of 0.5–3 μ m size. The 74% Fe₂O₃ sample contains sufficiently large particles, up to 40 μ m, which consist of sintered small particles, 0.5–1 μ m, while large particles are surrounded by a "grid" of small particles. The Ca/Fe ratio determined in the local surface areas ($d \approx 3-6$ μ m) of both samples is close to that in the ferrite phases, with a slight excess in Ca content in the 57% Fe₂O₃ sample.

The two-phase 35% Fe₂O₃ sample is represented by partially sintered particles of non-spherical, more elongated shape; these particles have size from 0.5 to 2–3 μ m. There is phase contrast in the images of this sample; according to EDX results the value of the Ca/Fe ratio in the bright areas is close to that for dicalcium ferrite with a slight excess in Ca content, and in the other areas it varies from 5 to 6. Another two-phase sample, 90% Fe₂O₃, consists of agglomerates of densely sintered particles of 1–3 μ m in size, surrounded by less sintered particles of similar sizes. The value of the Ca/Fe ratio in agglomerates is varied in the range of 0.01–0.1, and it is between 0.4 and 0.5 in the other parts of the sample.

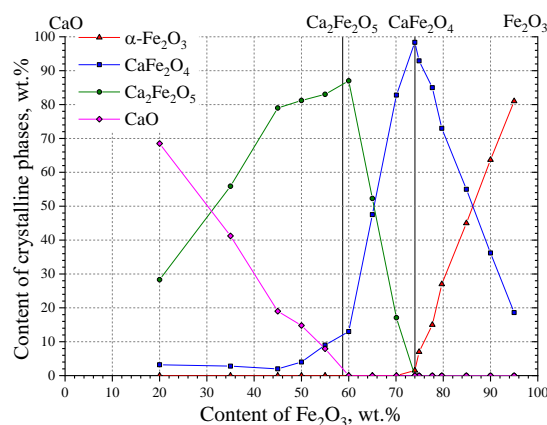


Figure 1 The catalysts phase composition versus the Fe₂O₃ content in the CaO-Fe₂O₃ system ($T_{synt.} = 900$ °C).

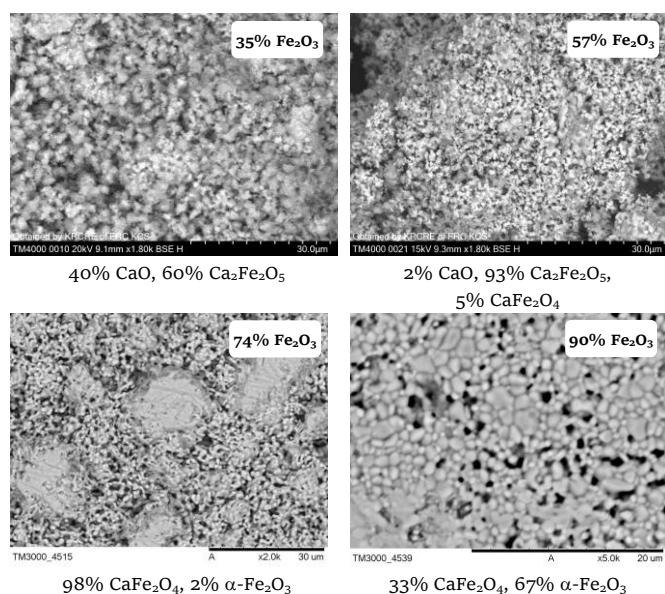


Figure 2 The SEM images of the CaO–Fe₂O₃ catalysts with different Fe₂O₃ content ($T_{\text{synt.}} = 1000\text{ °C}$): the catalyst particles (0.1–0.2 mm), flat surface of the 35 and 57% Fe₂O₃ samples, and the polished tablet surface of the 74 and 90% Fe₂O₃ samples.

One can see the EDS maps of these catalysts in Figure S4. The SEM–EDX study of catalysts synthesized at a lower temperature (900 °C) showed similar results for the samples of all compositions. It should be noted that the chemical composition, which was determined in the local surface areas of two-phase samples, does not correspond to the composition of the pure phases. In particular, it was shown that the formation of a core-shell structure is characteristic of the samples synthesized at 1000 °C with CaFe₂O₄– α -Fe₂O₃ compositions [14].

The effect of catalysts phase composition on the activity in the oxidative conversion of methane was investigated. CO₂, CO, C₂H₆ and H₂ are formed at the process conditions. It can be seen (Figure 3) that the rate of product formation in the presence of the catalysts synthesized at either 900 or 1000 °C varies with the Fe₂O₃ content.

In the region of two-phase CaO–Ca₂Fe₂O₅ samples ($T_{\text{synt.}} = 900\text{ °C}$), there is no additive change in activity and phase composition. The catalytic activity curve passes through a wide maximum in the range of 25–57 wt.% Fe₂O₃, and the composite catalysts CaO–Ca₂Fe₂O₅ with 35–55 wt.% Fe₂O₃ are the most active among all the catalysts. The patterns of changes in the specific activity of the active CaO–Ca₂Fe₂O₅ catalysts allow us to suppose that the active centers are localized at the interface of the phases CaO and Ca₂Fe₂O₅. It should be noted that stoichiometric ferrite phase Ca₂Fe₂O₅ undergoes a phase transition at temperatures $\sim 700\text{ °C}$, in which the symmetry of the unit cell changes from primitive to body-centered [15]. The XRD study at 750 °C of the two-phase samples CaO–Ca₂Fe₂O₅ showed that phase Ca₂Fe₂O₅ in the active catalysts is also in high-temperature modification at the catalytic temperature.

The activity of the Ca₂Fe₂O₅–CaFe₂O₄ catalysts decreases monotonically with an increase in the content of Fe₂O₃, and, accordingly, with a decrease in the Ca₂Fe₂O₅ content in the sample (Figure 1).

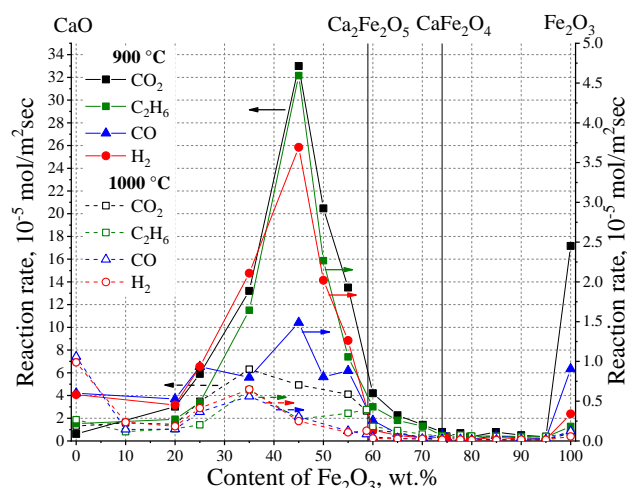


Figure 3 The dependences of the specific rate of product formation on the Fe₂O₃ content in the CaO–Fe₂O₃ system: filled point – $T_{\text{synt.}} = 900\text{ °C}$, empty characters – $T_{\text{synt.}} = 1000\text{ °C}$.

The two-phase catalysts of the CaFe₂O₄– α -Fe₂O₃ compositions have the least activity, while the formation rates of all products differ slightly and vary within small limits. The CO₂ and CO formation rates increase (as well as the formation rates of H₂ and C₂H₆ increase slightly) in the presence of single-phase hematite α -Fe₂O₃ sample, which belongs to the catalysts of deep oxidation of methane [16]. It is noteworthy that the activity of the CaFe₂O₄– α -Fe₂O₃ catalysts (74–95 wt.% Fe₂O₃) differs slightly regardless of the phase composition. As can be seen from the results of Figure 1, the α -Fe₂O₃ content in these catalysts varies in a wide range, from 1.5 to 81 wt.%. The appearance of the hematite phase in the catalyst composition and an increase in its content does not lead to an additive increase in the catalytic properties. Consequently, the low specific activity of CaFe₂O₄– α -Fe₂O₃ catalysts in the entire range of compositions is determined by the CaFe₂O₄ phase, which confirms the assumption of the core-shell structure formation in these catalysts, where the CaFe₂O₄ phase is a shell, based on the SEM–EDX results.

An increase in the synthesis temperature of catalysts to 1000 °C does not change the nature of the activity dependence (Figure 3), but leads to its decrease; at the points of maximum activity it is lower by more than 5 times. The high activity of low-temperature samples is apparently due to the formation of a more disordered Ca₂Fe₂O₅ phase with more mobile lattice oxygen, as evidenced by the large values of the b parameter of the unit cell of Ca₂Fe₂O₅ calculated from the XRD data. The b parameter varies in the ranges of 14.7830–14.7907 and 14.7759–14.7848 Å for the samples synthesized at 900 and 1000 °C, respectively. Ferrite Ca₂Fe₂O₅ has an orthorhombic structure of the brownmillerite-type, in which the octahedral layers alternate with the tetrahedral ones along the crystallographic axis b [15]. An increase in the values of parameter b indicates an elongation of Fe–O bonds, which leads to greater mobility of the apical oxygen (O_{ap}) atoms in octahedra's [17].

The selectivity of product formation (S) during the methane conversion on catalysts of the CaO–Fe₂O₃ oxide system is also different, which is well observed for the catalysts synthesized at 1000 °C (Figure 4). The most active catalysts of the CaO–Ca₂Fe₂O₅ compositions have high selectivity to CO₂ (S_{CO2}), 77–85%; S_{CO} and S_{C₂H₆} varies in the ranges of 3–8 and 10–16, respectively. An increase in the CaFe₂O₄ content in Ca₂Fe₂O₅–CaFe₂O₄ catalysts leads to a decrease in S_{CO2} and an increase in S_{C₂H₆}; values tend to selectivities of CaFe₂O₄–α-Fe₂O₃ catalysts, in the presence of which S_{CO2} and S_{C₂H₆} vary in ranges 43–54% and 40–58%. The catalysts synthesized at 900 °C have significantly different activity (Figure 3), while there is little difference in the selectivity of product formation (Figure S5). The S_{CO2} varies in the range of 71–82% for the catalysts with a content of 20–95% Fe₂O₃. The reason for the high selectivity to CO₂ of the CaFe₂O₄–α-Fe₂O₃ catalysts may be an increase in the contribution of the heterogeneous oxidation stage of both the CH₃-radical formed on the catalyst surface during activation of the methane molecule and the oxidation of the ethane molecule formed from two CH₃-radicals.

Thus, a study of the catalysts activity of the CaO–Fe₂O₃ oxide system ($T_{\text{synt.}} = 900$ and 1000 °C) in the oxidative conversion of methane at 750 °C showed that the most active catalysts are CaO–Ca₂Fe₂O₅ composite samples with a Fe₂O₃ content of 35–55 wt.%, which have a S_{CO2} of up to 85% and are more active compared with pure hematite, which is a catalyst of the deep methane oxidation.

4. Limitations

In high-temperature oxidation of hydrocarbons using complex oxide catalysts and molecular oxygen as an oxidizer, it is possible to proceed through the stages of selective and non-selective hydrocarbons conversion due to the participation of different active oxygen species/centers formed on the catalyst surface under reaction conditions. Establishing the real nature of the active centers is a difficult task that requires the use of experimental, including *in situ*, methods to study the surface state. We are going to study the catalysts' surface using the X-ray Photoelectron and Raman spectroscopy methods.

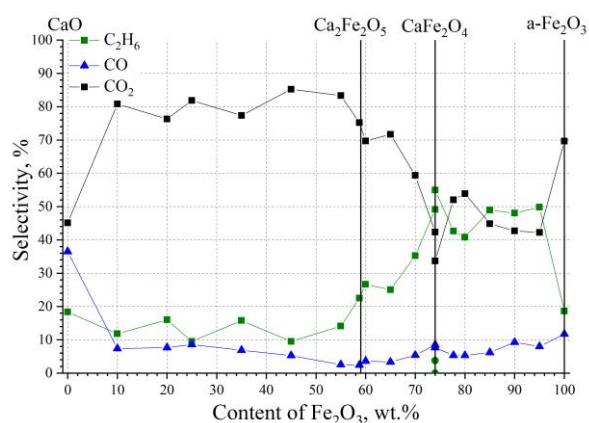


Figure 4 The dependences of products selectivity on the Fe₂O₃ content in the CaO–Fe₂O₃ system, $T_{\text{synt.}} = 1000$ °C.

5. Conclusions

The catalysts of the CaO–Fe₂O₃ oxide system with a variation of the Fe₂O₃ content from 0 to 100 wt.% were synthesized by the solid state method at 900 and 1000 °C. It was found that the specific activity of the catalysts in the oxidative conversion of methane at 750 °C varies and depends on the content of ferrite phases and their ratio.

The specific rate formation of C₂-hydrocarbons, CO and CO₂ in the presence of Ca₂Fe₂O₅–CaFe₂O₄ and CaFe₂O₄–α-Fe₂O₃ catalysts is lower compared with the catalysts based on CaO–Ca₂Fe₂O₅. The activity of the latter samples is due to the centers stabilized at the phase interface CaO/Ca₂Fe₂O₅. The ferrite phase Ca₂Fe₂O₅ has a high-temperature modification with a body-centered lattice at the catalysis temperature. The two-phase catalysts of the CaO–Ca₂Fe₂O₅ compositions with $T_{\text{synt.}} = 900$ and 1000 °C, have a high activity in deep oxidation of CH₄, while having a low, 10–16%, selectivity for C₂H₆.

The CaFe₂O₄–α-Fe₂O₃ catalysts with a core-shell structure possess low activity in the conversion of methane; the activity is determined by the shell that consists of CaFe₂O₄.

The synthesis of the samples at a higher temperature (1000 °C) leads to a decrease of the specific activity of all catalyst compositions, but increases the selectivity of ethane formation to 55% of the CaFe₂O₄–α-Fe₂O₃ catalysts.

The obtained results can be used in the development of catalytic systems based on calcium ferrites for high-temperature oxidation processes and chemical looping processes such as fuel combustion, coal and biomass gasification, hydrogen generation, etc.

• Supplementary materials

This manuscript contains supplementary materials, which are available on the corresponding online page.

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Methodology: N.K., E.R., L.S.

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● Conflict of interest

The authors declare no conflict of interest.

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